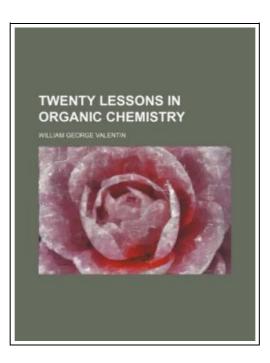
Twenty Lessons in Organic Chemistry



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Reviews

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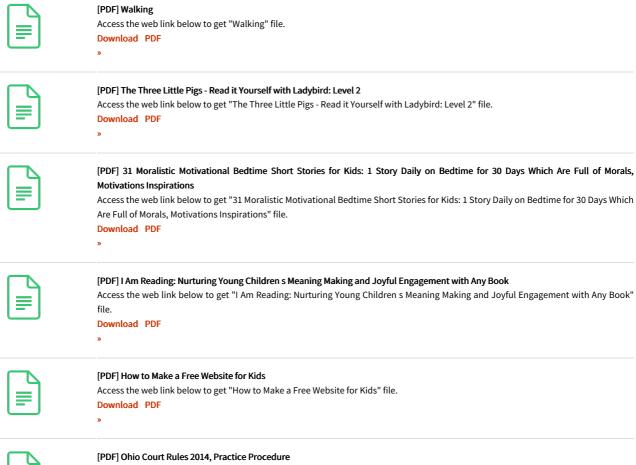
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Rarebooksclub.com, United States, 2012. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book ***** Print on Demand *****. This historic book may have numerous typos and missing text. Purchasers can download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1879 Excerpt: .10 0., and is generally sold in the form of sticks preserved under water, At the ordinary temperature it is soft and tough, at 0 C. it becomes brittle. It melts at 44 C, and is then of a yellowishsyrupy consistency, and boils at 288C. It volatilises, however, far below its boiling point, and as its vapour is spontaneously inflammable, the greatest care must be observed in handling this substance. Friction even may set it on fire. Phosphorus is insoluble in water, little soluble in alcohol and ether, very easily soluble, however, in carbon disulphide, from which solution it crystallises, out of contact of air, in rhombic dodecahedral crystals. This constitutes its first allotropic modification, viz., crystalline or yellow phosphorus. Its name is derived from the Greek word \$-trpipis, because it glows in the dark in air and other gaseous matter containing oxygen. This phenomenon appears, therefore, to be due to oxidation. When ordinary phosphorus is heated in iron vessels, with exclusion of air, to 300 C, for a few minutes even, it undergoes a change. The soft yellow phosphorus becomes a dark red and amorphous powder, and acquires materially different (physical) properties. The same change begins to take place slowly at 250 and upwards, and if a little iodine is added to the ordinary yellow phosphorus, it passes into another, the so-called red allotropic modification, even below 200 C. Beddish-brown amorphous phosphorus has a specific gravity of 2-14, and is insoluble in carbon disulphide, it...

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